

## Science 1206 - Naming Ionic Compounds.

The name of an ionic compound comes from the two ions (cation and anion) that make up the compound. **No prefixes** are used in naming ionic compounds.

Example:  $\text{MgCl}_2$  - Cation =  $\text{Mg}^{2+}$  = magnesium ion

Anion =  $\text{Cl}^-$  = chloride ion

Therefore:  $\text{MgCl}_2$  = magnesium chloride (not magnesuim dichloride)

Complete the table below:

Chemical Formula	Cation Symbol	Cation name	Anion Symbol	Anion name	Name of Compound
1. $\text{NaCl}$	$\text{Na}^+$	sodium	$\text{Cl}^-$	chloride	sodium chloride
2. $\text{CaO}$					
3. $\text{Ag}_2\text{S}$					
4. $\text{Al}_2\text{O}_3$					
5. $\text{K}_3\text{N}$					
6. $\text{LiBr}$					
7. $\text{AlP}$					
8. $\text{BaI}_2$					
9. $\text{Zn}_3\text{P}_2$					
10. $\text{Na}_2\text{O}$					
<p>Many anions are <u>polyatomic</u> - made up of 2 or more elements (See Table on reverse side of Periodic Table.) When polyatomic ions are involved, the rules for naming are the same.</p>					
11. $\text{BeSO}_4$	$\text{Be}^{2+}$	beryllium	$\text{SO}_4^{2-}$	sulfate	beryllium sulfate
12. $\text{Al}(\text{OH})_3$					
13. $\text{MgCO}_3$					
14. $\text{Sr}(\text{NO}_3)_2$					
15. $\text{Na}_2\text{SO}_3$					
16. $\text{KNO}_3$					
17. $\text{LiOH}$					
18. $\text{CaS}_2\text{O}_3$					