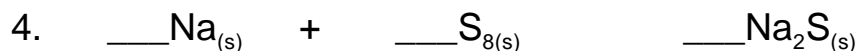
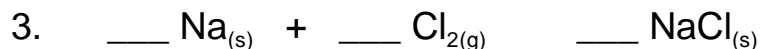
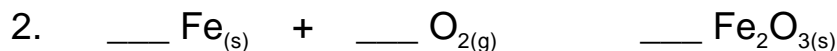
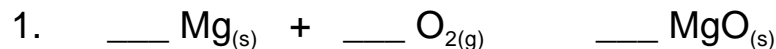


## Science 1206 - Balancing Equations and Reaction Types

The equations below are grouped by reaction type. Balance each equation:

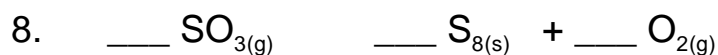
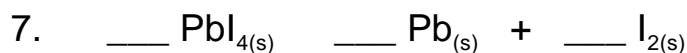
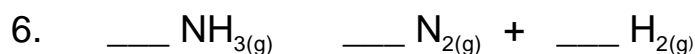
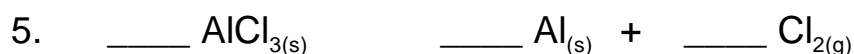
### A. Synthesis (Formation, Composition) -

2 elements (or 2 compounds) react to produce a single compound



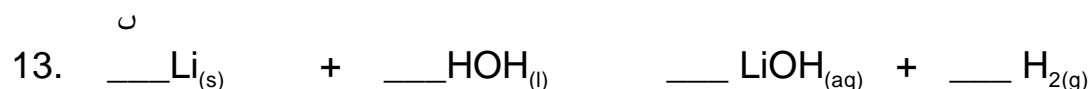
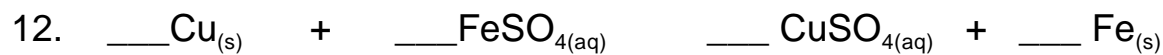
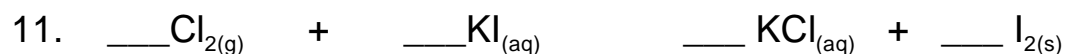
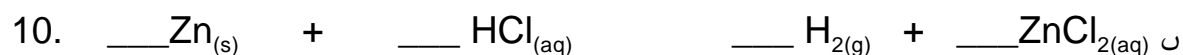
### B. Decomposition

A single compound is broken down (decomposed) into 2 or more elements &/or compounds



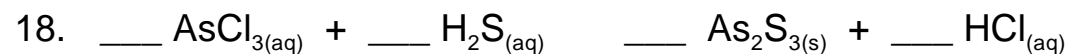
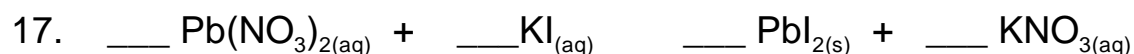
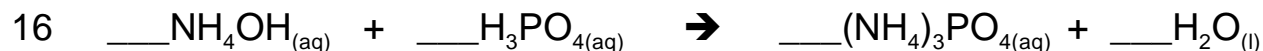
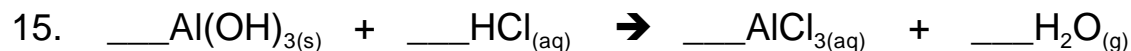
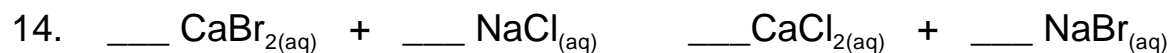
### C. Single Displacement (Single Replacement)

An element and a compound react to produce a new element and new compound



## D. Double Displacement (Double Replacement)

Usually 2 ionic compounds in aqueous solution are reacting  
Products are two compounds.



## E. Combustion

Oxygen reacts with a hydrocarbon to produce energy; reactant is "burned"  
If combustion is complete, carbon dioxide and water vapor are products

